

BAB V

KESIMPULAN DAN SARAN

V.1. Kesimpulan

Dalam penelitian ini, mempelajari pengaruh adsorpsi zat warna metilen biru dengan membandingkan hidrogel guar gum (HGG) dan HGG-TiO₂ komposit. Karakterisasi secara fisik dari hidrogel tersebut menggunakan analisa FTIR, XRD, SEM-EDX, dan BET. Proses adsorpsi yang dilakukan menggunakan *water bath shaker*, percobaan adsorpsi HGG tanpa menggunakan efek lampu UV, sedangkan pada percobaan HGG-TiO₂ menggunakan lampu UV untuk membantu proses photodegradasi warna metilen biru sehingga meningkatkan kapasitas adsorpsi. Isoterm adsorpsi menggunakan model persamaan Langmuir, Freundlich, Redlich-Peterson, Sips, dan Dubinin-Radushkevich, persamaan tersebut digunakan untuk mengkorelasikan data adsorpsi eksperimental, dari model persamaan isoterm adsorpsi yang digunakan menunjukkan proses adsorpsi yang terjadi secara monolayer. Pada penerapan kinetika adsorpsi terhadap waktu menunjukkan bahwa hasil yang terbaik pada pseudo orde satu. Hasil penerapan model termodinamika terhadap proses adsorpsi semakin naiknya suhu proses adsorpsi tidak spontan yang menandakan bahwa proses yang terjadi secara eksotermik.

V.2. Saran

Saran untuk penelitian berikutnya:

1. Proses adsorpsi terhadap zat warna lain oleh hidrogel guar gum dengan TiO₂ komposit perlu dipelajari lagi untuk mendapatkan kapasitas yang terbaik.

2. Perlu dipelajari lagi pengaruh ukuran hidrogel terhadap kapasitas adsorpsi yang terbaik.
3. Efek foto-degradasi terhadap metilen biru setelah penambahan TiO₂ perlu dipelajari.

DAFTAR PUSTAKA

- [1] A. T. Paulino, M. R. Guilherme, A. V. Reis, G. M. Campese, E. C. Muniz, and J. Nozaki, “Removal of methylene blue dye from an aqueous media using superabsorbent hydrogel supported on modified polysaccharide,” *J. Colloid Interface Sci.*, vol. 301, no. 1, pp. 55–62, 2006, doi: 10.1016/j.jcis.2006.04.036.
- [2] Menteri Negara Lingkungan Hidup, “Keputusan Menteri Negara Lingkungan Hidup Nomor: KEP-51/MENLH/10/1995 Tentang Baku Mutu Limbah Cair Bagi Kegiatan Industri,” *Menteri Negara Lingkung. Hidup*, 1995.
- [3] T. Hongbo, L. Yanping, S. Min, and W. Xiguang, “Preparation and property of crosslinking guar gum,” *Polym. J.*, vol. 44, no. 3, pp. 211–216, 2012, doi: 10.1038/pj.2011.117.
- [4] M. George and T. E. Abraham, “pH sensitive alginate-guar gum hydrogel for the controlled delivery of protein drugs,” *Int. J. Pharm.*, vol. 335, no. 1–2, pp. 123–129, 2007, doi: 10.1016/j.ijpharm.2006.11.009.
- [5] S. Thakur, B. Sharma, A. Verma, J. Chaudhary, S. Tamulevicius, and V. K. Thakur, “Recent approaches in guar gum hydrogel synthesis for water purification,” *Int. J. Polym. Anal. Charact.*, vol. 23, no. 7, pp. 621–632, 2018, doi: 10.1080/1023666X.2018.1488661.
- [6] W. Wang and A. Wang, “Preparation, swelling and water-retention properties of crosslinked superabsorbent hydrogels based on guar gum,” *Adv. Mater. Res.*, vol. 96, pp. 177–182, 2010, doi: 10.4028/www.scientific.net/AMR.96.177.
- [7] D. S. Warren, S. P. H. Sutherland, J. Y. Kao, G. R. Weal, and S. M. Mackay, “The preparation and simple analysis of a clay nanoparticle composite hydrogel,” *J. Chem. Educ.*, vol. 94, no. 11, pp. 1772–1779, 2017, doi: 10.1021/acs.jchemed.6b00389.
- [8] J. Chen, H. Park, and K. Park, “Synthesis of superporous hydrogels: Hydrogels with fast swelling and superabsorbent properties,” *J. Biomed. Mater. Res.*, vol. 44, no. 1, pp. 53–62, 1999, doi: 10.1002/(SICI)1097-4636(199901)44:1<53::AID-JBM6>3.0.CO.
- [9] E. Erizal, “The effect of hydrogel dressing copolymer poli(vinylpirrolidone) (pvp) - k- carrageenan prepared by radiation and healing times on the radius reductions burn injuried of wistar white rat,” *Indones. J. Chem.*, vol. 8, no. 2, pp. 271–278, 2010, doi: 10.22146/ijc.21633.
- [10] D. Jain, S. Verma, and S. Shukla, “Formulation and characterization

- of calcium chloride/guar gum microsphere of theophylline,” *Chronicles Young Sci.*, vol. 3, no. 2, p. 137, 2012, doi: 10.4103/2229-5186.98686.
- [11] E. S. Abdel-Halim and S. S. Al-Deyab, “Hydrogel from crosslinked polyacrylamide/guar gum graft copolymer for sorption of hexavalent chromium ion,” *Carbohydr. Polym.*, vol. 86, no. 3, pp. 1306–1312, 2011, doi: 10.1016/j.carbpol.2011.06.033.
- [12] S. Kaur and R. Jindal, “Synthesis of interpenetrating network hydrogel from (gum copal alcohols-collagen)-co-poly(acrylamide) and acrylic acid: Isotherms and kinetics study for removal of methylene blue dye from aqueous solution,” *Mater. Chem. Phys.*, vol. 220, pp. 75–86, 2018, doi: 10.1016/j.matchemphys.2018.08.008.
- [13] K. V. Kumar, “Linear and non-linear regression analysis for the sorption kinetics of methylene blue onto activated carbon,” *J. Hazard. Mater.*, vol. 137, no. 3, pp. 1538–1544, 2006, doi: 10.1016/j.jhazmat.2006.04.036.
- [14] I. A. W. Tan, A. L. Ahmad, and B. H. Hameed, “Adsorption of basic dye on high-surface-area activated carbon prepared from coconut husk: Equilibrium, kinetic and thermodynamic studies,” *J. Hazard. Mater.*, vol. 154, no. 1–3, pp. 337–346, 2008, doi: 10.1016/j.jhazmat.2007.10.031.
- [15] W. J. Thomas and B. Crittenden, *Adsorption Technology & Design*, Elsevier (1998).pdf, no. April. 1998.
- [16] D. D. Do, *Adsorption Analysis: Equilibria and Kinetics: (With CD Containing Computer Matlab Programs)*, vol. 2. 1998.
- [17] L. Laysandra *et al.*, “Highly adsorptive chitosan/saponin-bentonite composite film for removal of methyl orange and Cr(VI),” *Environ. Sci. Pollut. Res.*, vol. 26, no. 5, pp. 5020–5037, 2019, doi: 10.1007/s11356-018-4035-2.
- [18] O. Redlich and D. L. Peterson, “A Useful Adsorption Isotherm,” *J. Phys. Chem.*, vol. 63, no. 6, p. 1024, Jun. 1959, doi: 10.1021/j150576a611.
- [19] J. C. Y. Ng, W. H. Cheung, and G. McKay, “Equilibrium studies of the sorption of Cu(II) ions onto chitosan,” *J. Colloid Interface Sci.*, vol. 255, no. 1, pp. 64–74, 2002, doi: 10.1006/jcis.2002.8664.
- [20] F. Gimbert, N. Morin-Crini, F. Renault, P. M. Badot, and G. Crini, “Adsorption isotherm models for dye removal by cationized starch-based material in a single component system: Error analysis,” *J. Hazard. Mater.*, vol. 157, no. 1, pp. 34–46, 2008, doi: 10.1016/j.jhazmat.2007.12.072.
- [21] R. Sips, “On the structure of a catalyst surface,” *J. Chem. Phys.*, vol.

- 16, no. 5, pp. 490–495, 1948, doi: 10.1063/1.1746922.
- [22] A. Günay, E. Arslankaya, and I. Tosun, “Lead removal from aqueous solution by natural and pretreated clinoptilolite: Adsorption equilibrium and kinetics,” *J. Hazard. Mater.*, vol. 146, no. 1–2, pp. 362–371, 2007, doi: 10.1016/j.jhazmat.2006.12.034.
- [23] M. J. Ahmed and S. K. Dhedan, “Equilibrium isotherms and kinetics modeling of methylene blue adsorption on agricultural wastes-based activated carbons,” *Fluid Phase Equilib.*, vol. 317, pp. 9–14, 2012, doi: 10.1016/j.fluid.2011.12.026.
- [24] K. Y. Foo and B. H. Hameed, “Insights into the modeling of adsorption isotherm systems,” *Chem. Eng. J.*, vol. 156, no. 1, pp. 2–10, 2010, doi: 10.1016/j.cej.2009.09.013.
- [25] T. S. Anirudhan and P. S. Suchithra, “Heavy metals uptake from aqueous solutions and industrial wastewaters by humic acid-immobilized polymer/bentonite composite: Kinetics and equilibrium modeling,” *Chem. Eng. J.*, vol. 156, no. 1, pp. 146–156, 2010, doi: 10.1016/j.cej.2009.10.011.
- [26] H. Y., “Isotherms for the Sorption of Lead Onto Peat: Comparison of Linear and Non-Linear Methods,” *Polish J. Environ. Stud.*, vol. 15, no. 1, pp. 81–86, 2006.
- [27] D. A.O, “Langmuir, Freundlich, Temkin and Dubinin–Radushkevich Isotherms Studies of Equilibrium Sorption of Zn 2+ Unto Phosphoric Acid Modified Rice Husk,” *IOSR J. Appl. Chem.*, vol. 3, no. 1, pp. 38–45, 2012, doi: 10.9790/5736-0313845.
- [28] S. J. Allen, G. McKay, and J. F. Porter, “Adsorption isotherm models for basic dye adsorption by peat in single and binary component systems,” *J. Colloid Interface Sci.*, vol. 280, no. 2, pp. 322–333, 2004, doi: 10.1016/j.jcis.2004.08.078.
- [29] X. Chen, “Modeling of experimental adsorption isotherm data,” *Inf.*, vol. 6, no. 1, pp. 14–22, 2015, doi: 10.3390/info6010014.
- [30] S. P. Santoso *et al.*, “Eco-friendly cellulose–bentonite porous composite hydrogels for adsorptive removal of azo dye and soilless culture,” *Cellulose*, vol. 26, no. 5, pp. 3339–3358, 2019, doi: 10.1007/s10570-019-02314-2.
- [31] S. Hong, C. Wen, J. He, F. Gan, and Y. S. Ho, “Adsorption thermodynamics of Methylene Blue onto bentonite,” *J. Hazard. Mater.*, vol. 167, no. 1–3, pp. 630–633, 2009, doi: 10.1016/j.jhazmat.2009.01.014.
- [32] T. R. S. Cadaval, G. L. Dotto, and L. A. A. Pinto, “Equilibrium Isotherms, Thermodynamics, and Kinetic Studies for the Adsorption of Food Azo Dyes onto Chitosan Films,” *Chem. Eng. Commun.*, vol.

- 202, no. 10, pp. 1316–1323, 2015, doi: 10.1080/00986445.2014.934449.
- [33] S. K. Milonjić, “A consideration of the correct calculation of thermodynamic parameters of adsorption,” *J. Serbian Chem. Soc.*, vol. 72, no. 12, pp. 1363–1367, 2007, doi: 10.2298/JSC0712363M.
- [34] Y. S. Jeon, J. Lei, and J. H. Kim, “Dye adsorption characteristics of alginate/polyaspartate hydrogels,” *J. Ind. Eng. Chem.*, vol. 14, no. 6, pp. 726–731, 2008, doi: 10.1016/j.jiec.2008.07.007.
- [35] R. Sharma *et al.*, “Biodegradable and conducting hydrogels based on Guar gum polysaccharide for antibacterial and dye removal applications,” *J. Environ. Manage.*, vol. 162, pp. 37–45, 2015, doi: 10.1016/j.jenvman.2015.07.044.
- [36] V. K. Gupta, D. Pathania, P. Singh, A. Kumar, and B. S. Rathore, “Adsorptional removal of methylene blue by guar gum-cerium (IV) tungstate hybrid cationic exchanger,” *Carbohydr. Polym.*, vol. 101, no. 1, pp. 684–691, 2014, doi: 10.1016/j.carbpol.2013.09.092.
- [37] R. Sharma, S. Kalia, B. S. Kaith, and M. K. Srivastava, “Synthesis of guar gum-acrylic acid graft copolymers based biodegradable adsorbents for cationic dye removal,” *Int. J. Plast. Technol.*, vol. 20, no. 2, pp. 294–314, 2016, doi: 10.1007/s12588-016-9156-1.
- [38] J. N. Hiremath and B. Vishalakshi, “Evaluation of a pH-responsive guar gum-based hydrogel as adsorbent for cationic dyes: kinetic and modelling study,” *Polym. Bull.*, vol. 72, no. 12, pp. 3063–3081, 2015, doi: 10.1007/s00289-015-1453-x.
- [39] N. Thombare, U. Jha, S. Mishra, and M. Z. Siddiqui, “Borax cross-linked guar gum hydrogels as potential adsorbents for water purification,” *Carbohydr. Polym.*, vol. 168, pp. 274–281, 2017, doi: 10.1016/j.carbpol.2017.03.086.
- [40] T. Hongbo, L. Yanping, S. Min, and W. Xiguang, “Preparation and property of crosslinking guar gum,” *Polym. J.*, vol. 44, no. 3, pp. 211–216, 2012, doi: 10.1038/pj.2011.117.
- [41] X. Qin, A. Lu, and L. Zhang, “Gelation behavior of cellulose in NaOH/urea aqueous system via cross-linking,” *Cellulose*, vol. 20, no. 4, pp. 1669–1677, 2013, doi: 10.1007/s10570-013-9961-z.
- [42] T. K. Saha, N. C. Bhoumik, S. Karmaker, and M. G. Ahmed, “Adsorption of Methyl Orange onto Chitosan from Aqueous Solution,” vol. 2010, no. October, pp. 898–906, 2010, doi: 10.4236/jwarp.2010.210107.
- [43] S. Thakur and O. Arotiba, “Synthesis, characterization and adsorption studies of an acrylic acid-grafted sodium alginate-based TiO₂ hydrogel nanocomposite,” *Adsorpt. Sci. Technol.*, vol. 36, no.

- 1–2, pp. 458–477, 2018, doi: 10.1177/0263617417700636.
- [44] Merck, “IR Spectrum Table & Chart | Sigma-Aldrich,” *Sigmaaldrich*. 2020.
- [45] M. Grätzel, “Dye-sensitized solar cells,” *J. Photochem. Photobiol. C Photochem. Rev.*, vol. 4, no. 2, pp. 145–153, 2003, doi: 10.1016/S1389-5567(03)00026-1.
- [46] F. J. Sotomayor, K. A. Cybosz, and M. Thommes, “Characterization of Micro/Mesoporous Materials by Physisorption: Concepts and Case Studies,” *Acc. Mater. Surf. Res.*, vol. 3, no. 2, pp. 34–50, 2018.
- [47] M. Thommes *et al.*, “Physisorption of gases, with special reference to the evaluation of surface area and pore size distribution (IUPAC Technical Report),” *Pure Appl. Chem.*, vol. 87, no. 9–10, pp. 1051–1069, 2015, doi: 10.1515/pac-2014-1117.
- [48] W. Zhang, L. Zou, and L. Wang, “Photocatalytic TiO₂/adsorbent nanocomposites prepared via wet chemical impregnation for wastewater treatment: A review,” *Appl. Catal. A Gen.*, vol. 371, no. 1–2, pp. 1–9, 2009, doi: 10.1016/j.apcata.2009.09.038.
- [49] L. Laysandra *et al.*, “Adsorption and photocatalytic performance of bentonite-titanium dioxide composites for methylene blue and rhodamine B decoloration,” *Heliyon*, vol. 3, no. 12, p. e00488, 2017, doi: 10.1016/j.heliyon.2017.e00488.
- [50] W. Zhang, L. Zou, and L. Wang, “Visible-light assisted methylene blue (MB) removal by novel TiO₂/adsorbent nanocomposites,” *Water Sci. Technol.*, vol. 61, no. 11, pp. 2863–2871, 2010, doi: 10.2166/wst.2010.196.
- [51] J. J. M. Órfão *et al.*, “Adsorption of a reactive dye on chemically modified activated carbons - Influence of pH,” *J. Colloid Interface Sci.*, vol. 296, no. 2, pp. 480–489, 2006, doi: 10.1016/j.jcis.2005.09.063.
- [52] V. K. Agarwal, “Langmuir-Blodgett Films,” *Phys. Today*, vol. 41, no. 6, pp. 40–46, 1988, doi: 10.1063/1.881121.
- [53] L. Laysandra *et al.*, “Renewable rarasaponin-bentonite-alginate composite with sponge-like structure and its application for crystal violet removal from aqueous solution,” *Desalin. Water Treat.*, vol. 160, pp. 354–365, 2019, doi: 10.5004/dwt.2019.24196.
- [54] D. A. N. Yitzhaki, “Kinetics and Mechanism of Catalytic Oil : Adsorption and Hydrogenation Hydrodesulfurization of Gas of the Sulfur Compounds,” vol. 262, pp. 255–262, 1987.
- [55] J. Maity and S. K. Ray, “Enhanced adsorption of Cr(VI) from water by guar gum based composite hydrogels,” *Int. J. Biol. Macromol.*, vol. 89, no. Vi, pp. 246–255, 2016, doi:

- 10.1016/j.ijbiomac.2016.04.036.
- [56] L. Dai *et al.*, “Multifunctional self-assembling hydrogel from guar gum,” *Chem. Eng. J.*, vol. 330, pp. 1044–1051, 2017, doi: 10.1016/j.cej.2017.08.041.