

CHAPTER V

CONCLUSION AND RECOMMENDATION



CHAPTER V

CONCLUSION AND RECOMMENDATION

V.1 Conclusion

Jackfruit peel is suitable for activated carbon production using phosphoric acid activation for activation temperature at 450 and 550 °C. Surface area for adsorption (S_{BET}) of the activated carbons activated at 450 and 550 °C ranging from ± 900 – 1300 m²/g. Activation temperature 350 °C has not given pore development yet. N₂ adsorption isotherm and pore size distribution revealed that increasing impregnation ratio from 1 to 4 and activation temperature from 350, 450, and 550 °C gives more pore development to jackfruit peel. SEM images of the activated carbons showed more cavities of surface morphology of activated carbons as activation temperature increases while XRD patterns depicted that the activated carbons have graphite structure with typical of activated carbon.

Boehm titration, pH drift, and FTIR analyses show that activated carbons from jackfruit peel waste activated with phosphoric acid are rich in acidic surface functional groups. pH_{PZC} values of activated carbons are ranging from 1.9 – 2.0. Boehm titration result, confirmed by FTIR spectra, showed decreasing tendency of carboxylic and other acid group as activation temperature increases in the range of 350 °C – 550 °C. The acidic functional groups decreased after impregnation ratio 1 to 2, and with higher impregnation ratio from 2 to 3 and then 4, activated carbons tend to have slightly higher acidic groups.

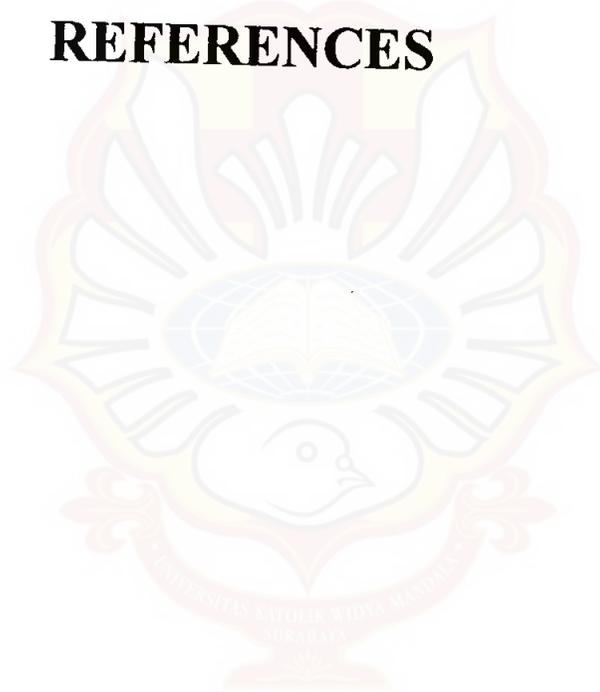
Jackfruit peel activated carbon was suitable on Methylene Blue adsorption. It was found that the jackfruit peel activated carbon with IR 4 and HTT 550°C had the highest removal capacity. IR[4]T[550] had a pH_{PZC} of 1.9 and was suitable for adsorption of Methylene Blue. With the highest initial pH solution (10.0) observed in this study, the activated carbon had enhancement both in kinetic and isotherm adsorption properties.

V.2 Recommendation

Recommendation for further experiment is as follows:

- Research of different alternatives of agricultural waste as precursor besides jackfruit peel is recommended.
- For jackfruit peel as precursor and phosphoric acid as activating agent, higher ratio of impregnation could be investigated in preparation of activated carbon. Wider range of activation temperature is also suggested. It seems that for jackfruit peel as precursor the optimum impregnation ratio and activation temperature may be different than usual optimum value given by literature. This was due to N_2 adsorption study which revealed that the highest value of those design variables still gave highest pore development.
- Further information about pore development progress of jackfruit peel could be carried out by means of analysis such as Thermogravimetric Analysis (TGA).
- For further identification, other methods of surface chemistry characterization such as X-Ray Photoelectron Spectroscopy (XPS) and Temperature Programmed Desorption (TPD) could be used.
- Another kind of adsorbate could be used to evaluate the activated carbon. Comparison of adsorption of two adsorbates, as example cationic dye and anionic dye, is also suggested.

REFERENCES



REFERENCES

- [1] Do, D. D., 1998, *Adsorption Analysis: Equilibria and Kinetics*, pp. 2-6. Imperial College Press, London.
- [2] Hu, Z. and Srinivasan, M.P., 1999, *Preparation of high-surface-area activated carbons from coconut shell*, *Microporous and Mesoporous Materials* 27, 11-18
- [3] Stavropoulos, G.G., Zabaniotou, A.A., 2005, *Production and characterization of activated carbons from olive-seed waste residue*, *Microporous and Mesoporous Materials*, 83, 79-85
- [4] Mozammel, H.M., Masahiro, O., Bhattacharya SC, 2002, *Activated charcoal from coconut shell using ZnCl₂ activation*, *Biomass and Bioenergy* 22, 397-400
- [5] Badan Pusat Statistik, 2006, *Selected Indicators Social-Economic of Indonesia*, p. 24, Directorate of Statistical Dissemination, Indonesia
- [6] Inbaraj, B.S., and Sulochana, N., 2003, *Carbonised jackfruit peel as an adsorbent for the removal of Cd(II) from aqueous solution*, *Bioresource Technology* 94, 49-52
- [7] Hu, Z., Srinivasan, M.P., Ni, Y., 2001, *Novel activation process for preparing highly microporous and mesoporous activated carbons*, *Carbon* 39, 877-886
- [8] Sudaryanto, Y., Hartono, S.B., Irawaty, W., Hindarso, H., Ismadji, S., 2006, *High surface area activated carbon prepared from cassava peel by chemical activation.*, *Bioresource Technology* 97, 734-739
- [9] Ahmadpour, A. and Do, D.D., 1996, *The preparation of active carbons from coal by chemical and physical activation.* *Carbon* 34, 471-479

- [10] Ahmadpour, A. and Do, D.D., 1997, *The preparation of activated carbon from macadamia nutshell by chemical activation*, Carbon 35, 1723-1732
- [11] Wibowo, N., Setiyadhi, L., Wibowo, D., Setiawan, J., Ismadji, S., 2007, *Adsorption of benzene and toluene from aqueous solutions onto activated carbon and its acid and heat treated forms: Influence of surface chemistry on adsorption*, Journal of Hazardous Material, Article in Press
- [12] El Qada, E.N., Allen, S.J., Walker, G.M., 2006, *Adsorption of Methylene Blue onto activated carbon produced from activated bituminous coal: A study of equilibrium adsorption isotherm*, Chemical Engineering Journal, 124, 103-110
- [13] Han, R., Zou, W., Yu, W., Cheng, S., Wang, Y., Shi, J., 2006, *Biosorption of methylene blue from aqueous solution by fallen phoenix tree's leaves*, Journal of Hazardous Materials 141, 156-162
- [14] Sekar, M., Sakthi, V., Rengaraj, S., 2004, *Kinetics and Equilibrium Adsorption Study of Lead (II) onto Activated Carbon Prepared from Coconut Shell*, Journal of Colloid and Interface Science 279, 307 – 313
- [15] IPTEKnet, 2005, *Teknologi Budaya Tanaman Pangan: Nangka Kunir*, in http://www.iptek.net.id/ind/teknologi_pangan/index.php?id=120 (Oktober 2006)
- [16] Maron, S. H. and Lando, J. B., 1974, *Fundamentals of Physical Chemistry*, p. 753-755, MacMillan Publishing Co Inc., New York
- [17] Smith, J.M., 1981, *Chemical Engineering Kinetics 3rd edition*, p. 311-324, Mc-Graw Hill, Singapore
- [18] El-Sheikh, A.H., Newman, A.P., Al-Daffaee, H.K., Phull, S., Cresswell, N., 2004, *Characterization of activated carbon prepared from a single cultivar of*

Jordanian Olive stones by chemical and physicochemical techniques, J. Anal. Appl. Pyrolysis 71, 151–164

- [19] Roy, G.M., 1995, *Activated Carbon Applications in the Food and Pharmaceutical Industries*, p. 1-14, Technomic Publishing co, Inc., Lancaster
- [20] McCabe, W.L. et al, 2001, *Unit Operations of Chemical Engineering 6th edition*, p. 506, 816-821, Mc-Graw Hill, Singapore
- [21] Tchobanoglous, G., 1991, *Wastewater Engineering: Treatment, Disposal, and Reuse, Third Edition*, Mc-Graw Hill, Inc., New York
- [22] Ismadji S., Sudaryanto, Y., Hartono, S.B., Setiawan, L.E.K., Ayucitra, A., 2005, *Activated carbon from char obtained from vacuum pyrolysis of teak sawdust: pore structure development and characterization*, Bioresource Technology 96, 1364–1369
- [23] Gómez-Serrano, V., Cuerda-Correa, E.M., Fernández-González, M.C., Alexandre-Franco, M.F., Macías-García, A., 2005, *Preparation of activated carbons from chestnut wood by phosphoric acid-chemical activation. Study of microporosity and fractal dimension*, Materials Letters 59, 846–853
- [24] Tseng, R.L. and Tseng, S.K., 2005, *Pore structure and adsorption performance of the KOH-activated carbons prepared from corncob*, Journal of Colloid and Interface Science 287, 428–437
- [25] Wu, F.C., Tseng, R.L., Hu, C.C., 2005, *Comparisons of pore properties and adsorption performance of KOH-activated and steam-activated carbons*, Microporous and Mesoporous Materials, 80, 95–106
- [26] Rodriguez-Reinoso, F., 1998, *The role of carbon materials in heterogeneous catalysts*, Carbon, 36, 159-175

- [27] Molina-Sabio, M. and Rodriguez-Reinoso, F., 2004, *Role of chemical activation in the development of carbon porosity*, Colloids and Surfaces A: Physicochem. Eng. Aspect, 241 (1-3), 15-25 sp
- [28] Srinivasakannan, C., and Bakar, M.Z.A., 2004, *Production of activated carbon from rubber wood sawdust*, Biomass and Bioenergy, 27, 89-96
- [29] Vernersson, T., Bonelli, P.R., Cerrella, E.G., Cukierman, A.L., 2001, *Arundo donax cane as a precursor for activated carbons preparation by phosphoric acid activation*, Bioresource Technology 83, 95-104
- [30] Bandosz T.J, Briggs M, Gubbins K.E., Kaneko K, Thomson K. Molecular models of porous carbons, 2003, In Chemistry and Physics of Carbon, Radovic L, ed. Marcel Dekker, New York, 98-103
- [31] Kadirvelu, K. and Namasivayam, C., 2002, *Activated carbon from coconut coirpith as metal adsorbent: adsorption of Cd (II) from aqueous solution*, Advances in Enviromental Research 7, 471-478
- [32] Chandra, T.C., Mirna, M.M, Sudaryanto, Y., Ismadji, S., 2006, *Adsorption of basic dye onto activated carbon prepared from durian shell: Studies of adsorption equilibrium and kinetics*. Chemical Engineering Journal 127, 121-129
- [33] Baquero, M.C., Giraldo, L., Moreno, J.C., Suárez-García, F., Martínez-Alonso, A., Tascón, J.M.D., 2003, *Activated carbons by pyrolysis of coffee bean husks in presence of phosphoric acid*, Journal of Analytical and Applied Pyrolysis 70, 779-784
- [34] Han, R., Wang, Y., Yu, W., Zou, W., Shi, J., Liu, H., 2006, *Biosorption of methylene blue from aqueous solution by rice husk in a fixed-bed column*, Journal of Hazardous Materials 141, 713-718

- [35] El-Sayed, Y. and Bandosz, T.J., 2004, *Adsorption of valeric acid from aqueous solution onto activated carbons: role of surface basic sites*, Journal of Colloid and Interface Science 273, 64-72
- [36] Figueiredo, J.L., Pereira, M.F.R., Freitas, M.M.A., Órfão, J.J.M., 1999, *Modification of the surface chemistry of activated carbon*, Carbon 39, 1379-1389
- [37] Diao, Y., Walawender, W.P., Fan, L.T., 2002, *Activated carbons prepared from phosphoric acid activation of grain sorghum*, Bioresource Technology 81, 45-52
- [38] Puziy, A.M., Poddubnaya, O.I., Martínez-Alonso, A., Suárez-García, F., Tascón, J.M.D., 2005, *Surface chemistry of phosphorus-containing carbons of lignocellulosic origin*, Carbon 43, 2857-2868
- [39] Guo, Y. and Rockstraw, D.A., 2006, *Physical and chemical properties of carbons synthesized from xylan, cellulose, and Kraft lignin by H₃PO₄ activation*, Carbon, 44, 1464-1475
- [40] Rodriguez-Reinoso, F. and Molina-Sabio, M., 1992, *Activated carbons from lignocellulosic materials by chemical and/or physical activation: an overview*, Carbon, 30, 1111-1118
- [41] Teng, H., Yeh, T.S., Hsu, L.H., 1998, *Preparation of activated carbon from bituminous coal with phosphoric acid activation*, Carbon, 36, 1387-1395
- [42] Jagtoyen, M., Derbyshire, F., 1998, *Activated carbons from yellow poplar and white oak by H₃PO₄ activation*, Carbon, 36, 1085-1097
- [43] Timur, S., Kantarli, I.C., Ikizoglu, E., Yanik, J., 2006, *Preparation of activated carbons from oreganum stalks by chemical activation*, Energy and Fuels, 20, 2636-2641

- [44] Hsu, L.Y., Teng, H., 2000, *Influence of different chemical reagents on the preparation of activated carbons from bituminous coal*, Fuel Processing Technology, 64, 155-166
- [45] Laine, J., Calafat, A., Labady, M., 1989, *Preparation and characterization of activated carbons from coconut shell impregnated with phosphoric acid*, Carbon, 27, 191-195
- [46] Ismadji, S. and Bhatia, S.K., 2001, *Characterization of Activated Carbons Using Liquid Phase Adsorption*, Carbon, 39, 1237-1250
- [47] Yang, T. and Lua, A.C., 2006, *Textural and chemical properties of zinc chloride activated carbons prepared from pistachio-nut shells*, Material Chemistry and Physics, 100, 438-444
- [48] Yang, R.T., 2003, *Adsorbent: Fundamentals and Applications*, pp. 86-89, John Wiley & sons Inc., New Jersey, USA
- [49] Boonamnuavitaya, V., Sae-ung, S., Tanthapanichakoon, W., 2003, *Preparation of activated carbons from coffee residue for the adsorption of formaldehyde*, Separation and Purification Technology 42, 159-168
- [50] Puziy, A.M., Poddubnaya, O.I., Martínez-Alonso, A., Suárez-García, F., Tascón, J.M.D., 2002, *Synthetic carbons activated with phosphoric acid I. Surface chemistry and ion binding properties*, Carbon 40, 1493-1505
- [51] Chingombe, P., Saha, B., Wakeman, R.J., 2005, *Surface modification and characterisation of a coal-based activated carbon*, Carbon 43, 3132-3143
- [52] Tor, A., Cengeloglu, Y., 2006, *Removal of congo red from aqueous solution by adsorption onto acid activated red mud*, Journal of Hazardous Materials B138, 409-415

- [53] ASTM International, 1999, *Standard Test Method for Total Ash Content of Activated Carbon*, **D 2866-94**, 813
- [54] ASTM International, *Standard Test Method for Moisture in Activated Carbon*, 1999, **D 2867-99**, 815-816
- [55] ASTM International, *Standard Test Method for Analysis of Volatile Matter Content of Activated Carbon Samples*, 2003, **D 5832-98**, 890-891

